

## MARK SCHEME for the May/June 2006 question paper

### 9701 CHEMISTRY

9701/02

Paper 2

Maximum raw mark 60

This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which Examiners were initially instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began. Any substantial changes to the mark scheme that arose from these discussions will be recorded in the published *Report on the Examination*.

All Examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes must be read in conjunction with the question papers and the *Report on the Examination*.

The minimum marks in these components needed for various grades were previously published with these mark schemes, but are now instead included in the Report on the Examination for this session.

- CIE will not enter into discussion or correspondence in connection with these mark schemes.

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Page 1	Mark Scheme	Syllabus	Paper
	GCE A/AS Level – May/June 2006	9701	02

- 1 (a) (i) ammonia/ $\text{NH}_3$  (1)
- (ii)  $\text{NH}_4^+$  (1)
- (iii) iron(II) hydroxide/ $\text{Fe}(\text{OH})_2$  (1) [3]
- (b) barium sulphate/ $\text{BaSO}_4$  (1) [1]
- (c) (i)  $\text{FeSO}_4$  (1)
- $(\text{NH}_4)_2\text{SO}_4$  (1)
- (ii)  $\text{FeSO}_4 = 151.9$  (allow 151.8 to 152) (1)
- $(\text{NH}_4)_2\text{SO}_4 = 132.1$  (allow 132) (1)
- (iii)  $x\text{H}_2\text{O} = 392 - (132.1 + 151.9) = 108$  (1)
- $x = \frac{108}{18} = 6$  (1)
- allow e.c.f. on candidate's sulphate in (c)(i)
- e.c.f. answer must be a whole number [6]

[Total: 10]

Page 2	Mark Scheme	Syllabus	Paper
	GCE A/AS Level – May/June 2006	9701	02

- 2 (a) 
$$\begin{array}{c} \text{x} \\ \text{o} \\ \text{H}^{\text{x}}\text{C}^{\text{x}}\text{C}^{\text{o}}\text{H} \\ \text{x} \\ \text{o} \end{array}$$
 (1) [1]
- (b)  $n = \frac{PV}{RT} = \frac{(1515 \times 10^3) \times (76 \times 10^{-3})}{8.31 \times 298}$  (1)
- $= 46.5$  (1) [2]
- (c) (i)  $\text{CaC}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$  (1)
- (ii)  $n(\text{C}_2\text{H}_2) = n(\text{CaC}_2) = 100 \times 46.5$  (1)
- mass of  $\text{CaC}_2 = 100 \times 46.5 \times 64 =$
- $= 297\,570 \text{ g}$
- $= 297.6 \text{ kg (accept 298 kg)}$
- correct units necessary (1)
- allow e.c.f. on candidate's answer in (b) [3]
- (d)  $\text{C}_2\text{H}_2(\text{g}) + \frac{5}{2}\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
- bonds broken:  $2(\text{H}-\text{C}) \quad 2 \times 410 \quad = \quad 820$
- $\text{C}\equiv\text{C} \quad 840 \quad = \quad 840$
- $\frac{5}{2}(\text{O}=\text{O}) \quad \frac{5}{2} \times 496 \quad = \quad \underline{1240}$
- $2900 \text{ kJ mol}^{-1}$  (1)
- bonds made:  $4(\text{C}=\text{O}) \quad 4 \times 740 \quad = \quad 2960$
- $2(\text{O}-\text{H}) \quad 2 \times 460 \quad = \quad \underline{920}$
- $3880 \text{ kJ mol}^{-1}$  (1)
- $\Delta H_{\text{comb}} = -3880 + 2900 = -980 \text{ kJ mol}^{-1}$  (1)
- allow e.c.f. on incorrect bonds made/broken [3]
- (e) (i) the enthalpy/energy change when one mole of a substance (1)
- is burned in an excess of air/oxygen
- or completely combusted
- under standard conditions (1)
- (ii) calculation in (d) includes  $\text{H}_2\text{O}(\text{g})$  whereas  $\Delta H_{\text{comb}}$  involves  $\text{H}_2\text{O}(\text{l})$
- or average bond energy terms are used in the *Data Booklet* (1) [3]

[Total: 12]

Page 3	Mark Scheme	Syllabus	Paper
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3 (a)

halogen	colour	physical state at room temperature
chlorine	green/yellow	gas
bromine	orange/red	liquid
iodine	black	solid

three colours correct (1)

three states correct (1) [2]

(b) (i)  $\text{MgCl}_2$  white fumes/steamy fumes (of  $\text{HCl}$ ) (1)

$\text{MgBr}_2$  red colour (of  $\text{Br}_2$ ) **or** steamy fumes (of  $\text{HBr}$ ) (1)

$\text{MgI}_2$  purple colour (of  $\text{I}_2$ ) **or**  
black solid ( $\text{I}_2$ ) **or**  
yellow solid (S) **or**  
stinking gas ( $\text{H}_2\text{S}$ ) (1)

(ii)  $\text{MgCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2\text{HCl}$   
allow  
 $\text{MgCl}_2 + 2\text{H}_2\text{SO}_4 \rightarrow \text{Mg}(\text{HSO}_4)_2 + 2\text{HCl}$  (1) [4]

(c) (i)  $\text{AgBr}$ /silver bromide (1)

(ii)  $\text{AgBr(s)} + 2\text{NH}_3(\text{aq}) \rightarrow \text{Ag}(\text{NH}_3)_2\text{Br(aq)}$   
equation (may be ionic) (1)

state symbols (1)

allow ecf on wrong halide in (i) [3]

(d) (i)  **$\text{HCl}$**  no reaction (1)

**$\text{HI}$**  purple vapour/black solid (1)

$2\text{HI} \rightarrow \text{H}_2 + \text{I}_2$  (1)

(ii) bond energy in  $\text{HCl}$  is high (1)

bond energy in  $\text{HI}$  is lower/more easily broken (1)

(iii) hot glass rod provides activation energy (1) [6]

[Total: 15]

Page 4	Mark Scheme	Syllabus	Paper
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- 4 (a) (i)  $\text{BrCH}_2\text{CHBrCH}_2\text{OH}$  (1)
- (ii)  $\text{CH}_2=\text{CHCO}_2\text{H}$  (1) [2]
- (b) oxidation (1) [1]
- (c) structural **or** functional group isomerism (1) [1]
- (d) **step I**  
 $\text{CH}_2=\text{CHCH}_2\text{OH} \rightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$   
 $\text{H}_2$  (1)  
 Ni catalyst and heat **or** Pt at room temperature (1)  
**step II**  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH} \rightarrow \text{CH}_3\text{CH}_2\text{CHO}$   
 acidified dichromate(VI) (1)  
 heat **not** under reflux  
**or** distil off aldehyde as it is formed (1)  
 for both steps, conditions mark is only awarded if correct reagent is given [4]
- (e) **both** oxidation and reduction have occurred (1) [1]
- (f) (i)  $\text{HOCH}_2\text{CHOHCH}_2\text{OH}$  formed  
 $\text{CH}_2=\text{CH}-$  forms  $\text{HOCH}_2\text{CHOH}-$  (1)  
 $-\text{CH}_2\text{OH}$  is unchanged (1)
- (ii)  $\text{HO}_2\text{CCO}_2\text{H}$  (1) [3]
- [Total: 12]

Page 5	Mark Scheme	Syllabus	Paper
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- 5 (a) A reaction in which one atom or group of atoms replaces another. (1) [1]
- (b) (i)  $\text{C}_2\text{H}_5\text{Br} + \text{NaOH} \rightarrow \text{C}_2\text{H}_5\text{OH} + \text{NaBr}$   
allow  $\text{OH}^-$  (1)
- (ii) heat with aqueous  $\text{NaOH}/\text{OH}^-$  (1) [2]
- (c) (i)  $\text{C}_2\text{H}_5\text{Br} + \text{KCN} \rightarrow \text{C}_2\text{H}_5\text{CN} + \text{KBr}$   
allow  $\text{CN}^-$  or  $\text{NaCN}$  (1)
- (ii) heat with  $\text{KCN}$  under reflux in ethanol (1) [2]
- (d) (i) K is  $\text{C}_2\text{H}_5\text{Cl}/\text{C}_2\text{H}_5\text{Br}/\text{C}_2\text{H}_5\text{I}$  (1)  
L is  $\text{C}_2\text{H}_5\text{CN}$  (1)
- (ii) **step I**  
reagent -  $\text{PCl}_3$  or  $\text{PCl}_5$  or  $\text{SOCl}_2$   
or  $\text{NaBr} + \text{conc H}_2\text{SO}_4$  or  $\text{P} + \text{Br}_2$   
or  $\text{P} + \text{I}_2$  (1)  
conditions - room temperature or heat under reflux depending on reagent (1)
- step III**  
reagent - mineral acid (1)  
conditions - heat under reflux (1)
- for both steps, conditions mark is only awarded if correct reagent is given [6]
- [Total: 11]